Amendments to the Specification:

Please replace paragraph [0044] with the following amended paragraph:

[0044] Figure 4 shows a graph of temperature of the solution and the hydrogen gas generation versus time for a $NaBH_4$ solution and a solution comprising $NaBH_4$ and 5% $LiBH_4$ by weight. The temperatures of the $NaBH_4$ solution and the $NaBH_4$/ $LiBH_4$ solution are indicated at ~~200~~ 201 and 204; corresponding ~~covers~~ curves for hydrogen flow rate, are indicated at 202 and 206 respectively. The temperature of the solution is directly related to the amount of heat given off during the exothermic reaction at a single point in time. Thus, the rate of temperature increase is an indirect way to monitor the reaction rate of the solution. Generally, the graph suggests that the $LiBH_4$-$NaBH_4$ solution has a slower rate of temperature increase, and thus a more stable hydrogen generation rate than the pure $NaBH_4$ solution. Accordingly, this makes the $LiBH_4$-$NaBH_4$ solution more desirable, as the hydrogen generation rate of this solution is more predictable and controllable. A hydrogen generation system must be capable of responding in real time to the fuel (hydrogen) needs of the fuel cell. This ability is referred to as the load following ability. Since the $LiBH_4$-$NaBH_4$ solution has a high hydrogen density and a controllable hydrogen generation rate, the solution has a good load following ability, and is ideally suited for this type of system.

Please replace paragraph [0050] with the following amended paragraph:

[0050] Still referring to Figure 7, there is also shown a hydrogen circulation loop or supplying path 500. Preferably, hydrogen is stored in the container 10 in the form of a gas. When hydrogen is required, the hydrogen gas is supplied to the first hydrogen connection port 110 through line 13 by means of a second pump 40. From here, the hydrogen enters the anode hydrogen inlet 140 (Figure 2), and flows along the channels in the flow field 132 on the front face of the anode flow field plate 120. At least a portion of the hydrogen diffuses across the first GDM 122 (Figure 1), and reacts at the anode

catalyst layer to form protons and electrons. The protons then migrate across the membrane 125 of the MEA 124 towards the cathode catalyst layer. The unreacted hydrogen continues to flow along the flow field 132, and ultimately exits the anode flow field plate 120 via the anode hydrogen outlet 141. From here, the unreacted hydrogen exits the stack via the second hydrogen connection port 111 (Figures 2a and 7), and returns to the container 10 via line 14.